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# Population ecology graph worksheet graph 2 sample design using excel

## Chapter 1 • Introduction

L1 A gas at 20°C may be rarefied if it contains less than  $10^{17}$  molecules per  $\text{mm}^3$ . If Avogadro's number is  $6.02 \times 10^{23}$  molecules per mole, what air pressure does this represent?

**Solution:** The mass of one molecule of air may be computed as  
$$m = \frac{\text{Molecular weight}}{\text{Avogadro's number}} = \frac{28.97 \text{ mol}^{-1}}{6.02 \times 10^{23} \text{ molecules/mol}} = 4.81 \times 10^{-26} \text{ g}$$
Then the density of air containing  $10^{17}$  molecules per  $\text{mm}^3$  is, in SI units,  
$$\rho = \left( 10^{17} \frac{\text{molecules}}{\text{mm}^3} \right) \left( 4.81 \times 10^{-26} \frac{\text{g}}{\text{molecule}} \right) = 4.81 \times 10^{-9} \frac{\text{g}}{\text{mm}^3} = 4.81 \times 10^{-6} \frac{\text{kg}}{\text{m}^3}$$

Finally, from the perfect gas law, Eq. (1.13), at 20°C = 293 K, we obtain the pressure:

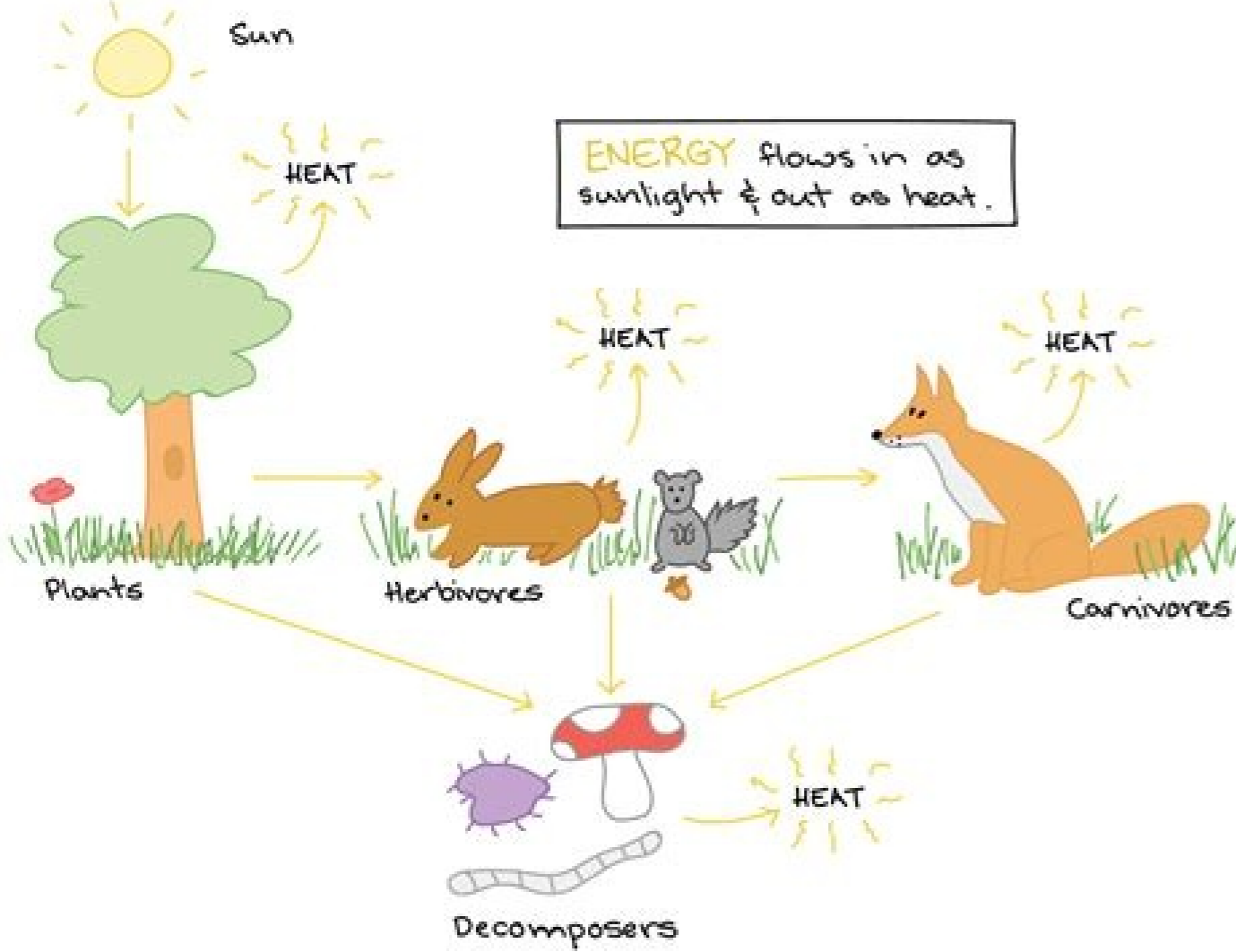
$$p = \rho RT = \left( 4.81 \times 10^{-6} \frac{\text{kg}}{\text{m}^3} \right) \left( 287 \frac{\text{m}^2}{\text{s}^2 \text{K}} \right) (293 \text{ K}) = 4.0 \text{ Pa} \quad \text{Ans.}$$

L2 The earth's atmosphere can be modeled as a uniform layer of air of thickness 20 km and average density  $1.2 \text{ kg/m}^3$  (see Table A-6). Use these values to estimate the total mass and total number of molecules of air in the entire atmosphere of the earth.

**Solution:** Let  $R_e$  be the earth's radius = 6377 km. Then the total mass of air in the atmosphere is  
$$m_a = \int \rho dV = \rho_{av} (\text{Air Vol}) = \rho_{av} 4\pi R_e^2 (\text{Air thickness}) = (1.2 \text{ kg/m}^3) (4\pi (6377 \text{ km})^2 (20 \text{ km})) = 6.1 \times 10^{18} \text{ kg} \quad \text{Ans.}$$

Dividing by the mass of one molecule =  $4.81 \times 10^{-26}$  g (see Prob. 11 above), we obtain the total number of molecules in the earth's atmosphere:

$$N_{\text{molecules}} = \frac{(\text{atmosphere})}{(\text{mass molecule})} = \frac{6.1 \times 10^{18} \text{ grams}}{4.81 \times 10^{-26} \text{ gm/molecule}} = 1.26 \times 10^{25} \text{ molecules} \quad \text{Ans.}$$



Excel is a powerful tool for data analysis and visualization. It allows you to create spreadsheets, charts, and graphs. You can use it to track your expenses, manage your budget, and analyze your data. Excel is used by millions of people around the world, and it is a valuable skill to have in many careers.

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When using Excel, it is important to be careful of your data. Make sure you are entering the correct information and that you are using the right formulas. It is also important to save your work regularly, so you don't lose any data. Excel is a powerful tool, but it can be tricky to use at first. Don't be afraid to ask for help if you need it.

Excel is a great tool for many different tasks. It can help you manage your budget, analyze your data, and create charts and graphs. If you are looking for a way to organize your data and make better decisions, Excel is a great choice. It is a powerful tool that can be used in many different ways. Whether you are a student, a professional, or a hobbyist, Excel can help you manage your data and make better decisions. If you are looking for a way to organize your data and analyze it, Excel is a great choice.

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